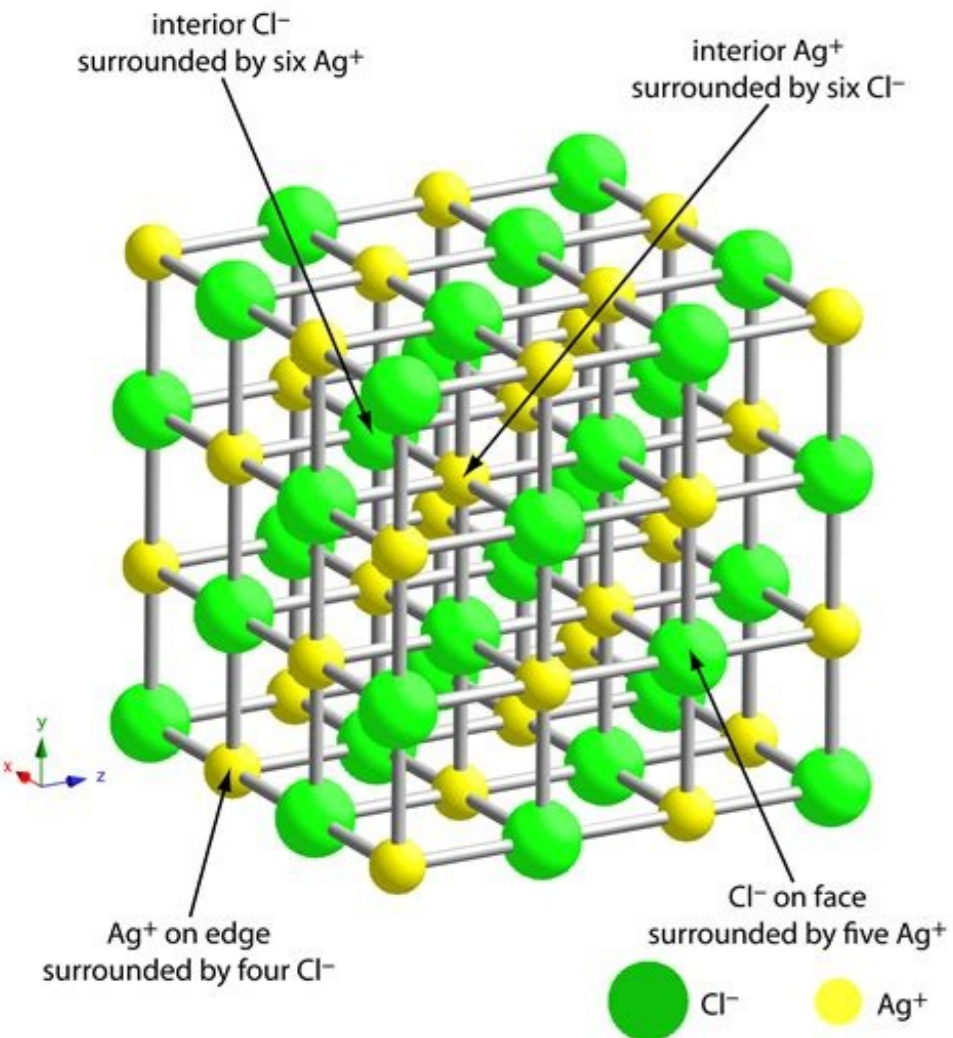


I'm not robot!



21- Use the following table and choose which of the species are neutral?

Atom or ion of element	A	B	C	D	E	F	G
Number of electrons (e <sup>-</sup> )	5	10	18	28	36	5	9
Number of protons (p)	5	7	19	30	35	5	9
Number of neutrons (n)	5	7	20	36	46	6	10

a- C and D    b- C, D, E, F and G    c- A, F and G    d- B and E

22- Use the above table and choose which of the species are negatively charged?

a- C and D    b- C, D, E, F and G    c- A, F and G    d- B and E

23- Use the above table and choose which of the species are positively charged?

a- C and D    b- C, D, E, F and G    c- A, F and G    d- B and E

24- Use the above table and choose which of the species have larger number of neutrons than protons?

a- C and D    b- C, D, E, F and G    c- A, F and G    d- B and E

25- The letter A in the chemical symbol ( ${}^A_ZX$ ) represents:

a- Atomic number    b- mass number    c- charge    d- frequency

26- The letter Z in the chemical symbol ( ${}^A_ZX$ ) represents:

a- Atomic number    b- mass number    c- charge    d- frequency

27- The letter c in the chemical symbol ( ${}^A_ZX^c$ ) represents:

a- Atomic number    b- mass number    c- charge    d- frequency

28- The letter f in the chemical symbol ( ${}^A_ZX^f$ ) represents:

a- Atomic number    b- mass number    c- charge    d- frequency

29- Which of the following is an isotope of  ${}^{208}_{82}\text{Pb}$ ?

a-  ${}^{207}_{81}\text{Tl}$     b-  ${}^{208}_{80}\text{Hg}$     c-  ${}^{212}_{82}\text{Pb}$     d-  ${}^{209}_{81}\text{Tl}$

30- Atoms of the same element with different mass numbers are called

a- ions.    b- neutrons.    c- isotopes.    d- chemical families.

3

Part A

Arrange the following substances in order of decreasing magnitude of lattice energy. Rank from largest to smallest magnitude of lattice energy. To rank items as equivalent, overlap them.

Reset

RbI    BaO

CaO    KCl

Largest magnitude    Smallest magnitude

**FORMS OF ENERGY BUNDLE**

- Potential/Kinetic Energy
- Examples Matching
- Forms of Energy Matching
- Energy Transfer Matching
- Identify Types of Energy in Common Objects Stations

Lattice energy  $\propto$  Total charge on ions

$\propto \frac{1}{\text{Size of ions}}$

① CsCl and BaCl<sub>2</sub>

CsCl  $\rightarrow$  Cs<sup>+</sup> + Cl<sup>-</sup>

BaCl<sub>2</sub>  $\rightarrow$  Ba<sup>2+</sup> + 2Cl<sup>-</sup>

**BaCl<sub>2</sub> > CsCl**

② RbI and RbBr

RbI  $\rightarrow$  Rb<sup>+</sup> + I<sup>-</sup>

RbBr  $\rightarrow$  Rb<sup>+</sup> + Br<sup>-</sup>

charge is same.

But size of I<sup>-</sup> is greater than Br<sup>-</sup>

**RbBr > RbI**

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